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Solution: OH OH OH 0.50 mmol mL ??

?=x = nc V ×25.0 mL OH n ? =12.5

mmol HCl(aq) HCl(aq) [HCl(aq)] 12.5

mmol = = n V 0.10 mmol HCl(aq) /mL

V =120 mL 2. (a) Given: V HCl(aq) =

25.00 mL; [HCl(aq)] = 0.350 mol/L; V

NaOH(aq) = 5.0 mL; [NaOH(aq)] =

0.500 mol/L; Required: [H+(aq)], pH

Analysis: Before titration, amount of

H+ is: H(aq) H(aq) H(aq) 0.350 mmol

mL ++ +=x = nc V

Section 8.7: Acid–Base Titration

Tutorial 1 Practice, page 547

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Solution: $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} = \frac{(0.040)(0.010)}{(0.14)^2} = 0.020$

Statement: The value of Q is 0.020. Q is equal to K , so the reaction is at equilibrium. (b) Given: $[HI(g)] = 0.20$

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mol/L; $[H_2(g)] = 0.15 \text{ mol/L}$; $[I_2(g)] = 0.090 \text{ mol/L}$; $K = 0.020$ Required: Q ;
direction of reaction to reach
equilibrium Solution: $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} = \frac{(0.15)(0.090)}{(0.20)^2}$
 $Q = 0.34$

Section 7.5: Quantitative Changes in Equilibrium Systems ...

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