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~~Reactions Unit 7~~

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~~AP Unit 7 AP~~

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~~molecular~~

~~compounds AP~~

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~~Chem 110 Unit 7~~

~~Organic~~

~~Chemistry~~

~~Segment 3~~

Chem 110 Unit 7

Segment 7**AP**

Chemistry Unit 7

FRQs Review Unit

7 Endothermic

reaction/graph

Chemistry,

class-12, unit-7

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~~Sheet Unit 7~~

are
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representative
of each type of
problem worked
in this unit.

For more
practice on a
given type of
problem, refer
back to Note
Taking Guides,
worksheets, and
quizzes. 1.

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Calculate the molar mass for ethane, C_2H_6 .

$$C_2H_6 \quad 2(12.0 \text{ g}) \\ + 6(1.0 \text{ g}) = \\ 30.0 \text{ g}$$

Convert 45 g Se to moles of Se.

$$? \text{ mol Se} = 45 \text{ g Se} \cdot \frac{1 \text{ mol Se}}{79.0 \text{ g Se}} = 0.57 \text{ mol Se}$$

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~~Unit 7 Name~~

~~Vocabulary: Fill
in the blanks~~

~~...~~

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of each type of
problem worked
in this unit.

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practice on a
given type of
problem, refer
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Taking Guides,
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quizzes. 1.

Calculate the
molar mass for
ethane, C_2H_6 .

2. Convert 45 g
Se to moles of
Se. 3. Convert
 4.77×10^{24}
molecules of SO_2
to grams.

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~~Unit 7 Name~~
~~Vocabulary: Fill~~
~~in the blanks~~

...

UNIT 7 REVIEW SHEET -

EQUILIBRIUM •
EQUILIBRIUM is
the point at
which the rate
of the forward
reaction is
equal to the
rate of the

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Reverse Reaction
in a reversible
process. o

Reaction is
dynamic - still
occurring - even
though to the
eye, it looks
like nothing is
changing at
equilibrium. o O
observable/measur
able quantities
(concentration,

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pressure, color)
are constant at
equilibrium. o
The amounts of
products and
reactants are
typically not
equal at
equilibrium!

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CHEMISTRY CP
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Name: _____ Per: _____

_____ TEST:

_____ Unit 7

Review Sheet:

Thermochemistry

KEY Potential

Energy Diagrams:

Choose the

appropriate

letter. 1. What

is the ΔH of the

forward

reaction? f 2.

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Review Sheet Unit 7 Answers

What is the activation energy of the forward

reaction? b 3.

What is the H of the reverse reaction? f 4.

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During the first
year of high
school chemistry

or the first

semester of

college

chemistry, all

the terms, units

of measurement,

and atoms,

molecules,

elements, and

compounds may

seem a bit

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overwhelming.
The good news:
Your Chem I
class doesn't
have to be
torture. This
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some basic
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Unit 5 Name _____

CHEMISTRY: A
Study of Matter

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5.24a KEY I. For each of the following statements, write "I" for ionic, "C" for covalent

~~CHEMISTRY: A~~

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Chemistry Unit 5

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2011 31. Balance

the following
equations by

putting the
proper

coefficients

wherever they

are needed. a)

$\text{Sr} + \text{O}_2 \rightarrow \text{SrO}$ b)

$\text{Al} + \text{H}_2\text{O}$

$\text{Al}(\text{OH})_3 + \text{H}_2$ c)

$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$

d) $\text{C}_5\text{H}_{12} + \text{O}_2$

$\text{CO}_2 + \text{H}_2\text{O}$ e)

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Na + H₂O → H₂ +
NaOH f) NaClO₃
NaCl + O₂ g ...

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