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Molecular Composition of Gases Chapter 11. What restrictions are there on the use.... At the same temperature and pressure, w.... According to Avogadro, what is the rela.... What is the relationship between the nu.... It applies only to gases measured under the same temperature a....

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334 CHAPTER 11 FIGURE 11-1 At the same temperature and pressure, balloons of equal volume have equal numbers of molecules, regardless of which gas they contain. Hydrogen molecule 1 mol H 2 at STP = 22.4 L Oxygen molecule 1 mol O 2 at STP = 22.4 L Carbon dioxide molecule 1 mol CO 2 at STP = 22.4 L

CHAPTER 11 Molecular Composition of Gases

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Chapter 11 - Molecular Composition of Gases I. Measuring and Comparing the Volumes of Reacting Gases A. Observations of Gay-Lussac 1. 2 liters H 2 + 1 liter O 2 \u2262 2 liters H 2O vapor 2. 2 volumes H 2 + 1 volume O 2 \u2262 2 volumes H 2O vapor 3. 1 volume H 2 + 1 volume Cl 2 \u2262 2 volumes HCl 4.

Chapter 11 - Molecular Composition of Gases

CHAPTER 11 REVIEW Molecular Composition of Gases MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided 1 c The average speed of a gas molecule is most directly related to the (a) polarity of the molecule (b) pressure of the gas (c) temperature of the gas (d) Chapter 11 Review Gases Mixed Answer Key

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Chapter 11 Review Molecular Composition Of Gases Mixed Answers Answer. Molar mass = 151.1 g/mol Potassium: (39.10 / 101.1) x 100 = 38.67% Nitrogen: (14.01 / 101.1) x 100 = 13.86% Oxygen: (48.00 / 101.1) x 100 = 12.58%.

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CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided 1 The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol (c) multiplied by 224 L. (b) divided by the mass of 1 mol (d) divided by 224 L 2 For the expression V nR P T